

Table 2 Effect of solvents on $\text{CuCl}_2/\text{AlEt}_2\text{Cl}$ catalyst system for styrene polymerization

Solvent	ϵ	% Conversion	Reaction time (hours)
n-Hexane	1.89	38.8	0.5
Benzene	2.28	43.0	0.5
Toluene	2.37	36.5	0.5
Bromobenzene	5.40	29.7	0.5
Chlorobenzene	5.62	24.4	0.5
$\text{C}_2\text{H}_4\text{Cl}_2$	10.36	47.9	0.5

$\text{CuCl}_2 + \text{AlEt}_2\text{Br}$ catalyst system for the polymerization of styrene². The rate does not become constant after adding the specific amount of PPh_3 , as was seen previously². Hence the conclusion can be drawn, that PPh_3 forms a complex with copper chloride and inhibits the polymerization.

Ethers as additives were studied in detail by Cunningham⁴ and Gippin⁵. Different products have been identified in the reactions of ether with Ziegler-Natta type catalysts. The effect of these products mainly govern the reaction rate. The effect of anisole can be explained on the basis of its reaction with AlEt_2Cl which may give various products which retard the polymerization. Gippin⁶ observed no change in the rate at optimum catalyst ratio when diphenylether (DPE) was added as an additive. The

present study also confirms the same observations.

Pyridine and triethylamine totally inhibit the polymerization as in the case of $\text{CuCl}_2 + \text{AlEt}_2\text{Br}$ system². Hydroquinone also did not effect the rate indicating absence of free radicals.

As can be seen from Table 2, the results obtained in solvents of varying dielectric constant indicate a decrease in conversion with increase in dielectric constant. The halogenated aliphatic solvents again indicate higher yields than expected on the basis of their dielectric constants. This has also been observed and explained earlier².

CONCLUSIONS

Above studies indicate that the mechanism of polymerization of styrene by $\text{CuCl}_2 + \text{AlEt}_2\text{Cl}$ catalyst system is of the anionic coordinated type, though the activation energy is low.

REFERENCES

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